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| Name:  |   | <br>UNIVERSITY OF PETROLEUM & ENERGY STUDIES |            |
| Enrolment No:  |   |   |            |
| <b>UPES</b><br><b>End Semester Examination, May 2023</b> |   |   |            |
| <b>Programme Name:</b> B.Tech. (APEG)                    |   | <b>Semester</b> : VI  |            |
| <b>Course Name</b> : Unconventional Gas Resources        |   | <b>Time</b> : 3 Hrs.  |            |
| <b>Course Code</b> : CHGS3002P                           |   | <b>Max. Marks</b> : 100   |            |
| <b>Nos. of page(s)</b> : 2                               |   |   |            |
| <b>Instructions:</b>                                     |   |   |            |
| 1. All questions are compulsory.                         |   |   |            |
| 2. Assume any missing data, if any.                      |   |   |            |
| <b>S. No.</b>  | <b>Section - A (4Qx5M = 20)</b>   | <b>Marks</b>  | <b>CO</b>  |
| <b>Q1</b>  | Differentiate between face cleat and butt cleat in a coal matrix.   | <b>4</b>  | <b>CO1</b> |
| <b>Q2</b>  | Differentiate between conventional reservoir and unconventional reservoir.  | <b>4</b>  | <b>CO1</b> |
| <b>Q3</b>  | List the factors that can be identified by logging in tight reservoirs.   | <b>4</b>  | <b>CO3</b> |
| <b>Q4</b>  | Describe the normal ways of gas storage in gas shale.   | <b>4</b>  | <b>CO1</b> |
| <b>Q5</b>  | Enumerate the characteristics of shale gas.   | <b>4</b>  | <b>CO2</b> |
| <b>Section - B (4Qx10M = 40)</b>                         |   |   |            |
| <b>Q6</b>  | A 0.7 gravity gas is at 600 psia and 90 <sup>0</sup> F. To what value can the temperature be reduced without hydrate formation. Use hydrate equilibrium curve given in appendix 1.  | <b>10</b>   | <b>CO4</b> |
| <b>Q7</b>  | Describe the characteristics of the various gas hydrate structures.   | <b>10</b>   | <b>CO4</b> |
| <b>Q8</b>  | Discuss the additives commonly used in fracturing fluid and their functions.  | <b>10</b>   | <b>CO3</b> |
| <b>Q9</b>  | Explain the process of fracturing in tight reservoir with the help of schematic diagram.  | <b>10</b>   | <b>CO3</b> |
| <b>Section - C (2Qx20M = 40)</b>                         |   |   |            |
| <b>Q10</b>   | Calculate C, H, N, S % from the following observations for a sample of coal:<br>a) 2.1 g of coal is burnt in combustion tube. The increase in weight of anhydrous CaCl <sub>2</sub> is 0.53 gm and increase in weight of KOH tube is 5.73 gm.<br>b) 0.75 gm of coal in Kjeldahl's experiment released NH <sub>3</sub> , which is passed in 50 ml 0.12 N HCl. The HCl requires 39 ml of 0.12 N NaOH to neutralize in back titration. | <b>20</b>   | <b>CO4</b> |

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|            | c) Washings of bomb pot when 1.9 gm of the coal sample in bomb calorimeter experiment is treated with BaCl <sub>2</sub> solution to give 0.41 gm BaSO <sub>4</sub>   |           |            |
| <b>Q11</b> | Discuss the following:<br>a) Gas Shale vs Shale Gas (Marks - 6)<br>b) Langmuir adsorption Isotherm vs Freundlich adsorption isotherm (Marks - 7)<br>c) Industry challenges associated with exploration and exploitation of shale gas (Marks - 7) | <b>20</b> | <b>CO3</b> |

**Appendix 1 - Hydrate Equilibrium Curve**

