

Structural and Functional Studies on Two-Dimensional Ternary Coordination Polymers

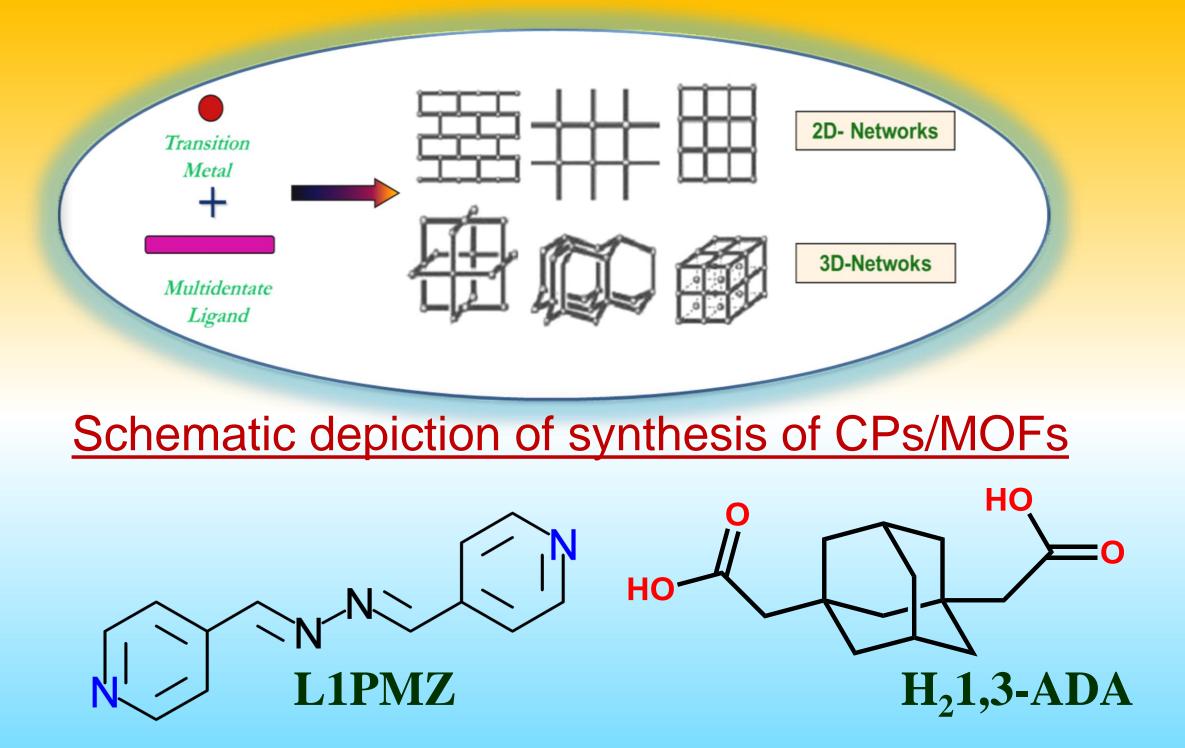
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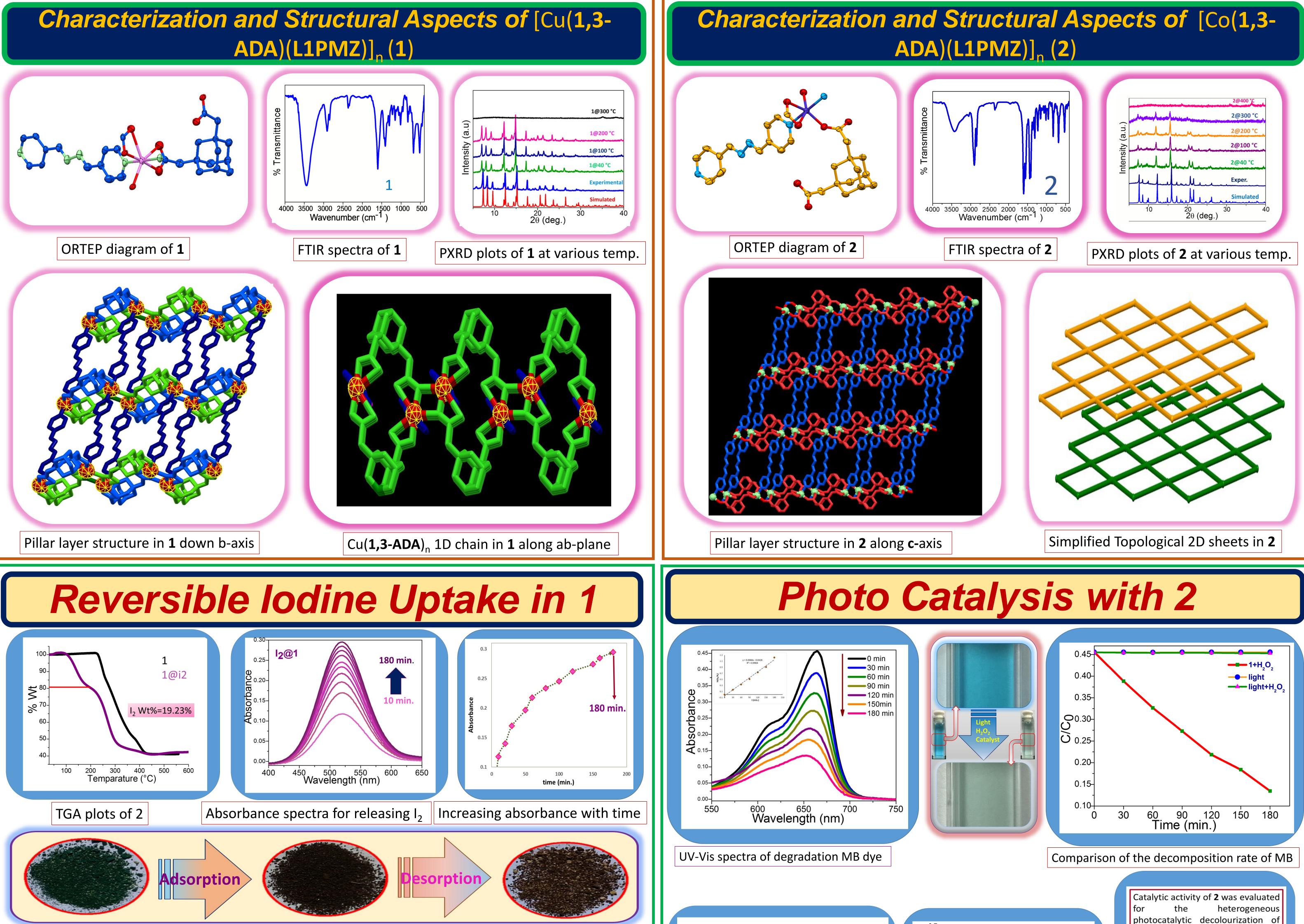
Coordination Polymers(CPs) or Metal-Organic Frameworks (MOFs) are extended polymeric solid materials which are being explored for last two decades for the applications in numerous fields including sensing and photocatalysis due to their modular (tunable) syntheses, porosity and large surface area.

As an extension of our endeavors to synthesize novel CPs, herein, we have chosen semirigid linear ligand L1PMZ (N,N-bis-pyridin-4-ylmethylene-hydrazine) along with flexible dicarboxylate 1,3-ADA (1,3-adamantane diacetate) as the spacers and metal



nodes Cu^{2+}/Co^{2+} to synthesize new CPs $[Cu(1,3-ADA)(L1PMZ)]_n$ (1), and $[Co(1,3-K)]_n$ **ADA**)(L1PMZ)]_n (2). Crystal structure of 1 and 2 revealed interesting 2D structures. These materials have shown functional capabilities towards iodine adsorption and photocatalytic degradation of methyleneblue (MB) dye.

Ligands explored in present work



20 (deg.) 20

Digital Photographs of Iodine encapsulated Samples

Iodine adsorption study was carried out for **1**. 40 mg of **1** was taken in small vial and placed in closed containers and exposed to the iodine vapor at 80°C for 6 4 h. Releasing property was studied in hexane with time dependent UV-Vis spectra recorded at an interval of 10 min for 3 h presents the absorbance (λ_{max}) 524 nm) enhancement due to increasing iodine concentration.

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Comparison PXRD patterns of 1

After Catalysis Enuction -2.0 -2.0 -2.0 -2.0 **Experimental** K-⊼ Simulated Band Gap Energy 2.07 eV 2.0 2.5 3.0 3.5 4.0 4.5 5.0 5.5 6.0 Energy (eV) 2θ (deg.)

Comparison PXRD patterns of **2**

methylene blue (MB) dye by 2-3 drops of 30% H₂O₂ solution at ambient temperature. In the typical experiment, 100 mL glass reactor was charged with 50 mL aqueous MB solution $(1 \times 10^{-5} \text{ M})$ and the 15 mg photocatalyst (2) was dispersed The reaction was illuminated by a tungsten filament lamp. Stirring was maintained throughout the reaction and 1 mL aliquot was withdrawn every 30 min which was then centrifuzed and analyzed by absorption spectrophotometer.

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References:

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Band gap energy plot of **2**